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CARBON AND ITS COMPOUNDS

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[Carbon And Its Compounds](#)

Carbon and its compounds is a really interesting study. Let us do a comparison between a diamond and charcoal. One is a shiny, attractive, hard rock, and the other a black, ashy, soft substance. But both are elements of carbon! Their varied properties are only due to the different arrangement of the carbon atoms.

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Chemical Properties of Carbon Compounds: The important chemical properties are as follows: 1. Combustion: The complete combustion of carbon compounds in the air gives carbon dioxide water, heat and light. $\text{CH}_3\text{CH}_2\text{OH(l)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O(l)} + \text{Heat and light}$. Carbon burns in air or oxygen to give carbon dioxide and heat and light.

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Friendly Carbon Why Carbon Can Form so Many Compounds. Catenation occurs most readily with carbon due to its small size, electronic configuration and unique strength of carbon-carbon bonds. Tetravalency, catenation and tendency to form multiple bonds with other atoms account for the formation of innumerable carbon compounds.

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Compounds of carbon are defined as chemical substances containing carbon. More compounds of carbon exist than any other chemical element except for hydrogen. Organic carbon compounds are far more numerous than inorganic carbon compounds. In general bonds of carbon with other elements are covalent bonds.

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Abundance: Carbon is the 4th most abundant substance in universe and 15th most abundant substance in the earth's crust. Compounds having carbon atoms among the components are known as carbon compounds. Previously, carbon compounds could only be obtained from a living source; hence they are also known as organic compounds.

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Carbon compounds generally burn (oxidize) in air to produce carbon dioxide and water, and release heat and light energy. Saturated hydrocarbon burns generally with a blue flame in good supply of air and with a yellow sooty flame in limited supply of air. Sooty flame is seen when unsaturated hydrocarbons are burnt.

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5) Formation of a very large number of carbon compounds :- Carbon forms a very large number of compounds. The number of carbon compounds is more than three million. It is more than the number of...

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12. Chemical Properties of Carbon Compounds : (i) Combustion : Carbon compound undergo combustion reaction to produce CO_2 and H_2O with the evolution of heat and light. $\text{CH}_4 + \text{O}_2 > \text{CO}_2 + 2\text{O} + \text{heat and light}$ (ii) Oxidation : The substance which are used for oxidation are known as oxidising agent. e.g alkaline KMnO_4 , acidified $\text{K}_2\text{Cr}_2\text{O}_7$.

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Carbon (from Latin: carbo "coal") is a chemical element with the symbol C and atomic number 6. It is nonmetallic and tetravalent—making four electrons available to form covalent chemical bonds. It belongs to group 14 of the periodic table. Carbon makes up only about 0.025 percent of Earth's crust.

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1) Occurrence of carbon :- i) Carbon is found in the atmosphere, inside the earth's crust and in all living organisms. ii) Carbon is present in fuels like wood, coal, charcoal, coke, petroleum, natural gas, biogas, marsh gas etc. iii) Carbon is present in compounds like carbonates, hydrogen carbonates etc. iv) Carbon is found in the free state as diamond, graphite, fullerenes etc.

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[Some Important Carbon Compounds: Types, Videos and Examples](#)

Carbon and its Compounds. 1. Carbon is found in the form of compounds except in diamond, graphite, coal etc. which are obtained freely in nature. Some of the main sources of carbon are as follows: a. In air, carbon dioxide is found. b. Various petroleum products. c. In all living organisms, in substances like carbohydrate, starch, urea etc. 2. Artificial sources: methane, alcohol, glycerol ...

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Carbon and its Compounds Extra Questions Very Short Answer Type. Question 1. Draw the electron dot structure of the gas molecule which is liberated when zinc metal is treated with aqueous NaOH solution. Answer: Hydrogen gas, Question 2. Write the number of covalent bonds in the molecule of ethane. Answer: There are seven covalent bonds : Question 3. Write the number of covalent bonds in the ...

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Carbon plays very important roles for all living beings. The amount of carbon in the earth's crust is merely 0.02%, which is available in the form of minerals such as carbonates, hydrogen-carbonates, coal, and petroleum. The presence of carbon in the atmosphere of the earth is 0.03%, in the form of carbon dioxide. Compounds of Carbon

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Selina Concise ICSE Solutions for Class 8 Chemistry Chapter 9 Carbon and Its Compounds. Points to Remember: Carbon occurs in the earth's crust in the free as well as in the combined state. In the free state, it occurs as coal, diamond and graphite. In the combined state, carbon occurs in atmosphere (CO_2) natural gas, food nutrients and carbonates. Diamond is the hardest naturally occurring ...

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Carbon, chemical element that forms more compounds than all the other elements combined. Carbon is widely distributed in coal and in the compounds that make up petroleum, natural gas, and plant and animal tissue. The carbon cycle is one of the most important of all biological processes.

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Ans. Carbon and its compounds undergo combustion to produce heat, the amount of heat released can be handled and used so they are used as fuels for most applications. Q10. Explain the formation of scum when hard water is treated with soap. Ans. Hard water contains salts of calcium and magnesium. When soap molecule comes in contact with these salts it forms a curdy white precipitate (compound ...

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Carbon and Its Compounds Organic Compounds. Initially, compounds of carbon could only be obtained from living sources and there was no way of synthesizing them. Hence, carbon compounds are also known as organic compounds. Carbon forms a large number of compounds. So far, formulae of about 3 million carbon compounds are known. Cause of formation of such a large number of compounds by carbon ...

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• Carbon and its compounds are used as fuels because they burn in air releasing lot of heat energy. • Saturated hydrocarbon generally burn in air with blue and non-sooty flame. • Unsaturated hydrocarbon which does not get yellow sooty flame because percentage of carbon is higher than saturated hydrocarbon which does not get completely oxidized in air. (ii) Oxidation. Alcohols can be ...

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