

MODERN CHEMISTRY CHAPTER 5 SECTION 3 REVIEW ANSWERS

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a. 2 and 4 c. 4 and 5 b. 3 and 4 d. 5 and 6 ____ 21. The discovery of the noble gases changed Mendeleev's periodic table by adding a new. a. period. c. group. b. series. d. sublevel block. ____ 22. In the modern periodic table, elements are ordered according to. a. decreasing atomic mass. c. increasing atomic number. b. Mendeleev's original ...

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

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CHAPTER 1 REVIEW Matter and Change SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. A horizontal row of elements in the periodic table is called a(n) period . 2. The symbol for the element in Period 2, Group 13, is B . 3. Elements that are good conductors of heat and electricity are metals . 4.

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Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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(Section 1.4) Report the answers to calculations to the correct number of significant figures. (Section 2.2) Make unit conversions. (Section 2.5) 3.1 Solids, Liquids, and Gases 3.2 The Chemical Elements 3.3 The Periodic Table of the Elements 3.4 The Structure of the Elements 3.5 Common Elements 3.6 Relating Mass to Number of Particles. 76 Chapter 3 The Structure of Matter and the Chemical ...

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(Section 5.3) Report the answers to calculations to the correct number of significant figures. (Section 8.2) Use percentages as conversion factors. (Section 8.4) Make unit conversions. (Section 8.5) All of these people ask questions that begin, "How much...?" Review Skills The presentation of information in this chapter assumes that you can already perform the tasks listed below. You can ...

[mc06se cFMsr i-vi](#)

CHAPTER . 12 . REVIEW ""-- Solutions . SHORT ANSWER Answer the following questions in the space provided. 1. Describe the errors made by the following students in making molar solutions. 3. James needs a 0.600 M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the solid. James made several errors. First, 0.600 mol of KCl does not have a mass of 0.600 g. Also, adding ...

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Chapter Quiz Class 5.2 5.3 5.1 5.1 5.3 5.3 5.3 5.3 5.3 5.3 115 Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT. The orbitals of a principal energy level are lower in energy than the orbitals in the next higher principal energy level. 3. In the quantum mechanical model of the atom, the probability of

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CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: H₂SO₃(aq) + H₂O(l) ~ H₃O⁺(aq) + HSO₃⁻(aq) b. Which stage of ionization usually produces more ions? Explain your answer.

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